

Predicting Redox Products

For each of the following reactions, write a BALANCED equation, coefficients should be in terms of lowest whole numbers. Assume that solutions are aqueous unless otherwise indicated. Represent substances in solutions as ions if the substances are extensively ionized. **Omit formulas for any ions or molecules that are unchanged by the reaction.**

Magnesium ribbon is burned in oxygen.

Acidified solutions of potassium permanganate and iron (II) nitrate are mixed together.

A piece of copper wire is placed in a solution of silver nitrate.

Chlorine gas is bubbled into a solution of magnesium bromide.

Bromine gas is bubbled into a solution of dilute sodium hydroxide.

A solution of potassium permanganate is mixed with an alkaline solution of sodium sulfite.

Acidified potassium permanganate solution is added to an acidic solution of hydrogen peroxide.